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# Determination of TCAP (Thermal Cycling Absorption Process) Raffinate Purity

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*36<sup>th</sup> Tritium Focus Group Meeting*

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# Analytical Issue and Solution in H<sub>2</sub>/D<sub>2</sub> R&D

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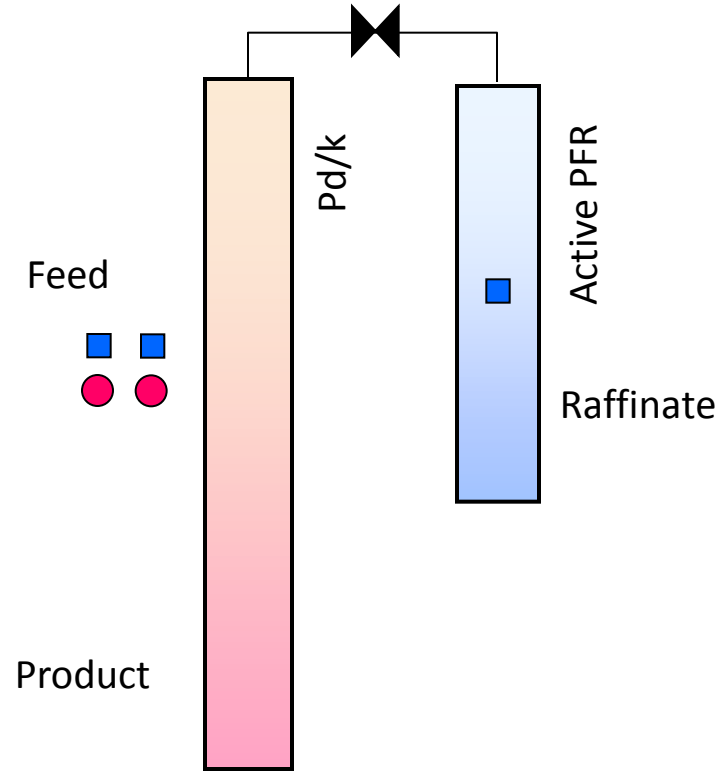
- 99+% D<sub>2</sub> (product) and H<sub>2</sub> (raffinate) from TCAP, better purity than detection limit
- Raffinate has much less D<sub>2</sub> than ultra high purity gas cylinder (natural abundance D/H ratio is 155.78 ppm).
- Raffinate impurity correlates with H<sub>2</sub> stack emission in tritium operation

- ~ 200 sl TCAP raffinate gas was collected in a hydride bed
- The hydride bed was transferred to another room
- The raffinate in hydride bed was oxidized into water in a catalyst evaluation lab reactor
- D/H ratio was analyzed by spectroscopy method



# TCAP Working Principle

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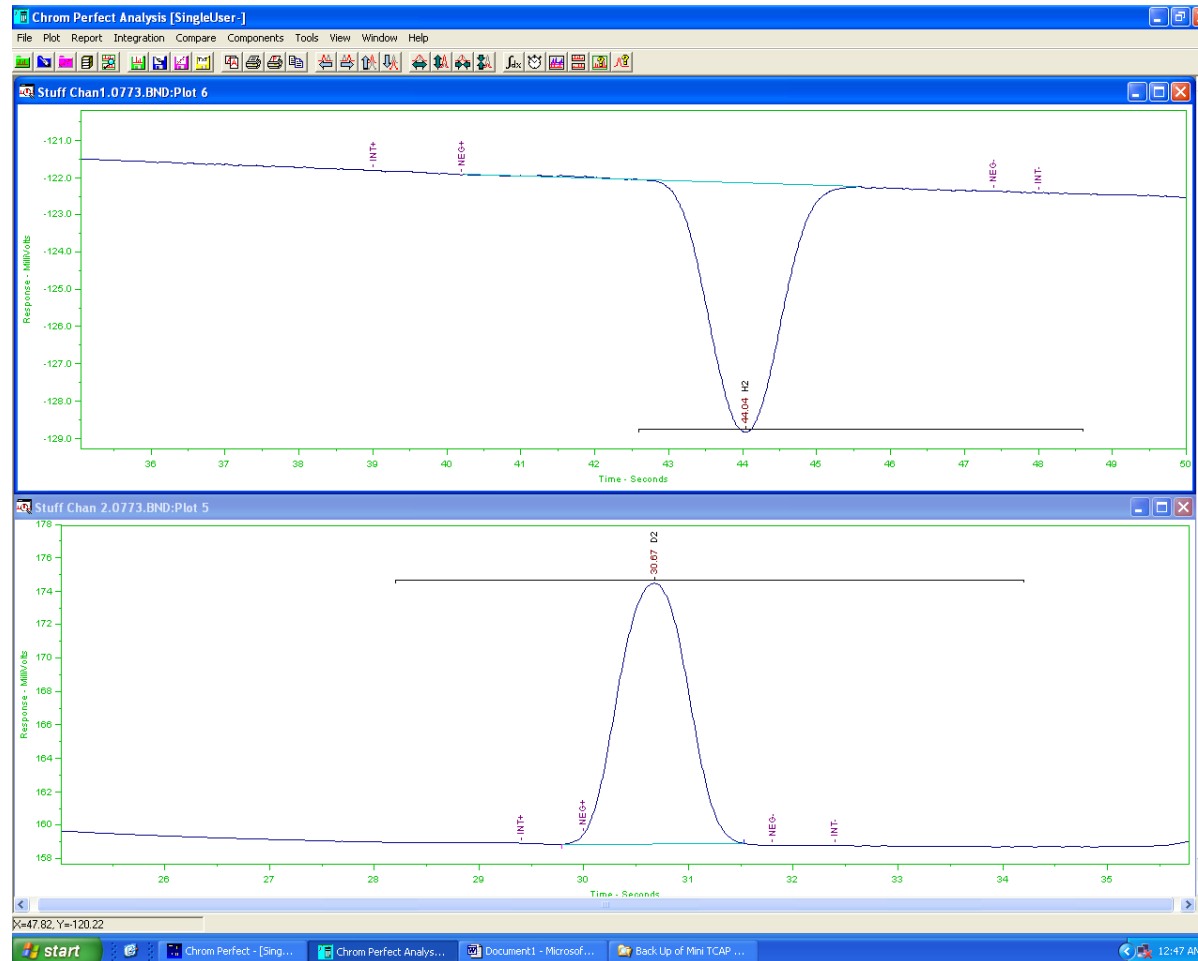


# Previous Analysis Using Thermal Conductivity (TCD)

H<sub>2</sub> analysis  
using D<sub>2</sub> carrier

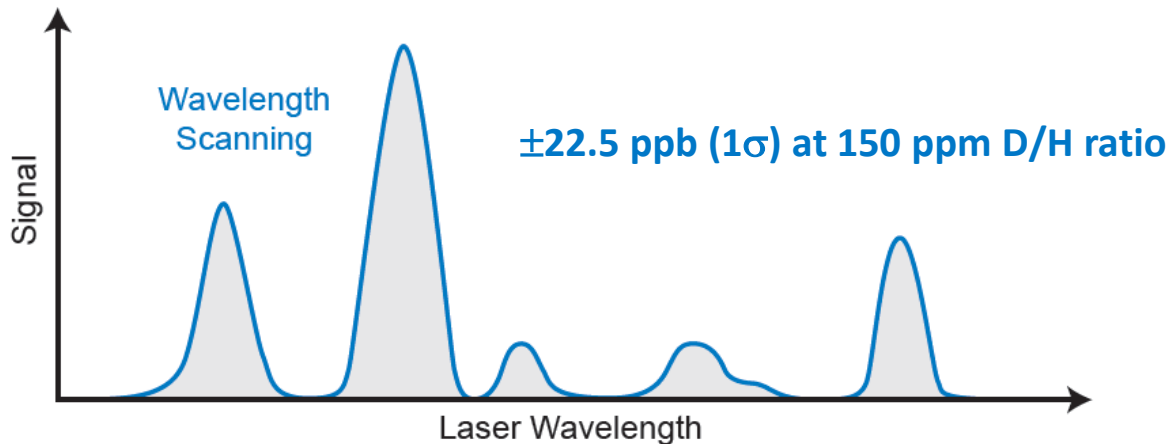
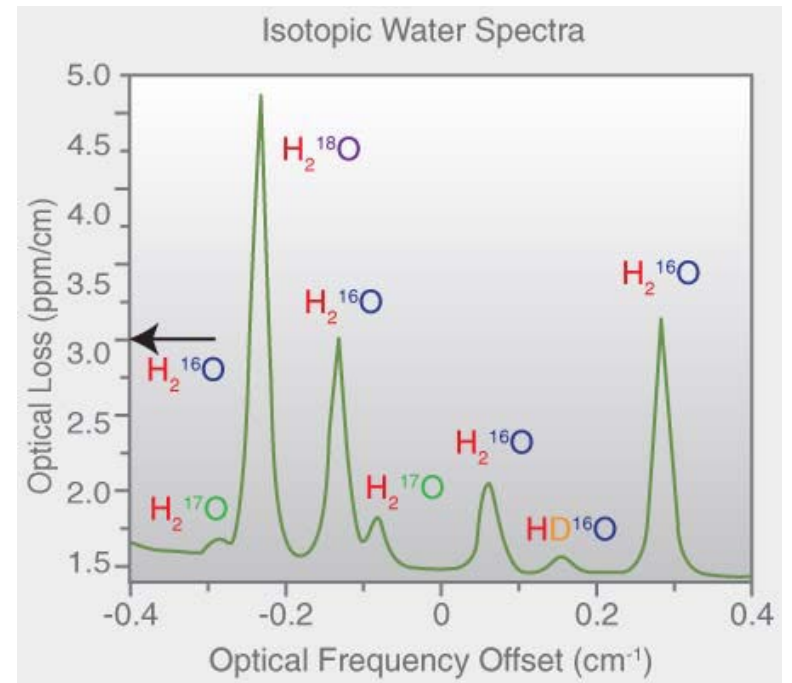
10 ppm resolution  
lack of low range  
calibration standards

D<sub>2</sub> analysis  
using H<sub>2</sub> carrier



*H. T. Sessions Jr., "Analytical Method for Measuring Total Protium and Total Deuterium in a Gas Mixture Containing H<sub>2</sub>, D<sub>2</sub>, and HD via Gas Chromatography", presented at Tririum Conference 2007, Rochester, NY*

# Tunable Laser for Water Isotope Analysis



# Pathways Considered to Oxidize Hydrogen into Water, Safely!

Reaction	Est. Avg. $\Delta T$	$T_{Max}$ (adiabatic)
$H_2 + CuO \xrightarrow{550\text{ K}} H_2O + Cu$	20.8°C	1790°C
$Cu + 1/2 O_2 (Air) \xrightarrow{550\text{ K}} CuO$	15.5°C	1890°C
$H_2 + 1/2 O_2 (Air) \xrightarrow[550\text{ K}]{Pt/Al_2O_3} H_2O$	57.8°C	3042°C
$H_2 + 1/2 O_2 (Air) \xrightarrow[ambient]{Pt/PSDB} H_2O$	67.6°C	2790°C

**Hydrogen autoignition: 565°C**

**Reactor Design Max Temperature: 538°C**



# Reduction of Cupric Oxide by Hydrogen

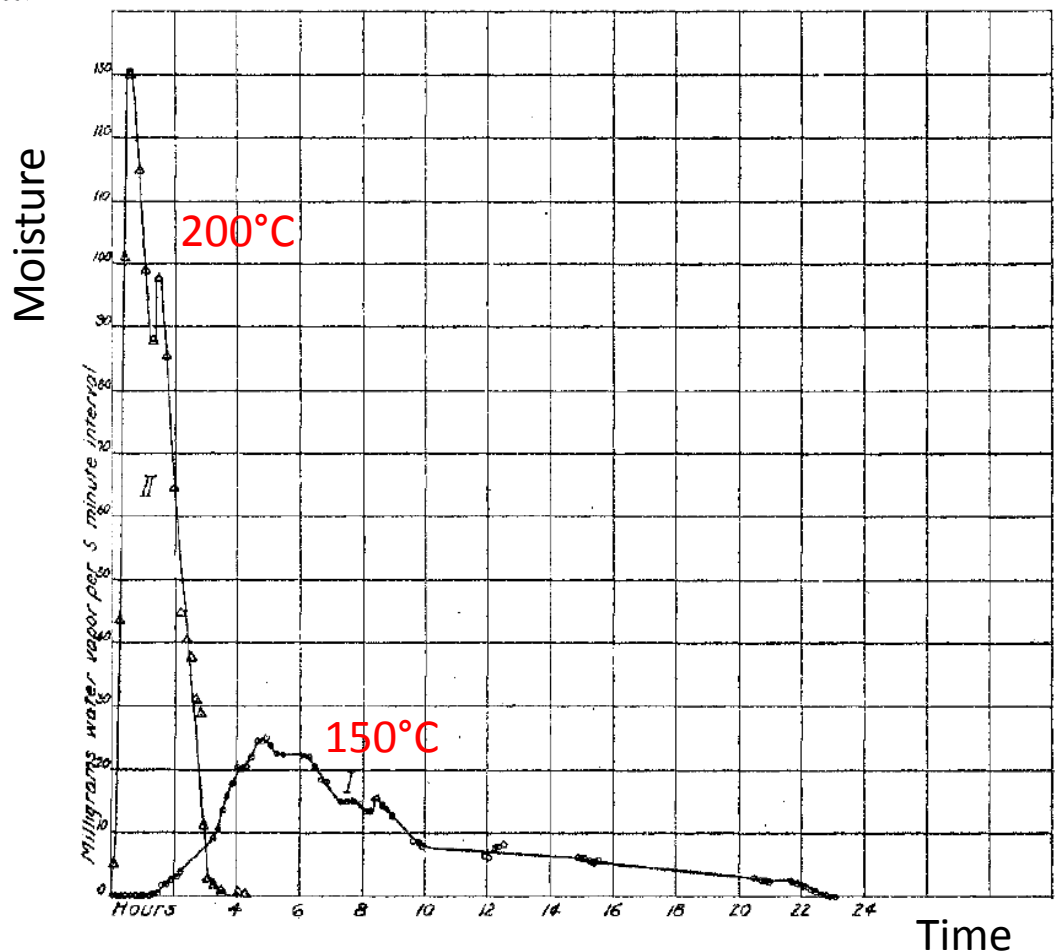
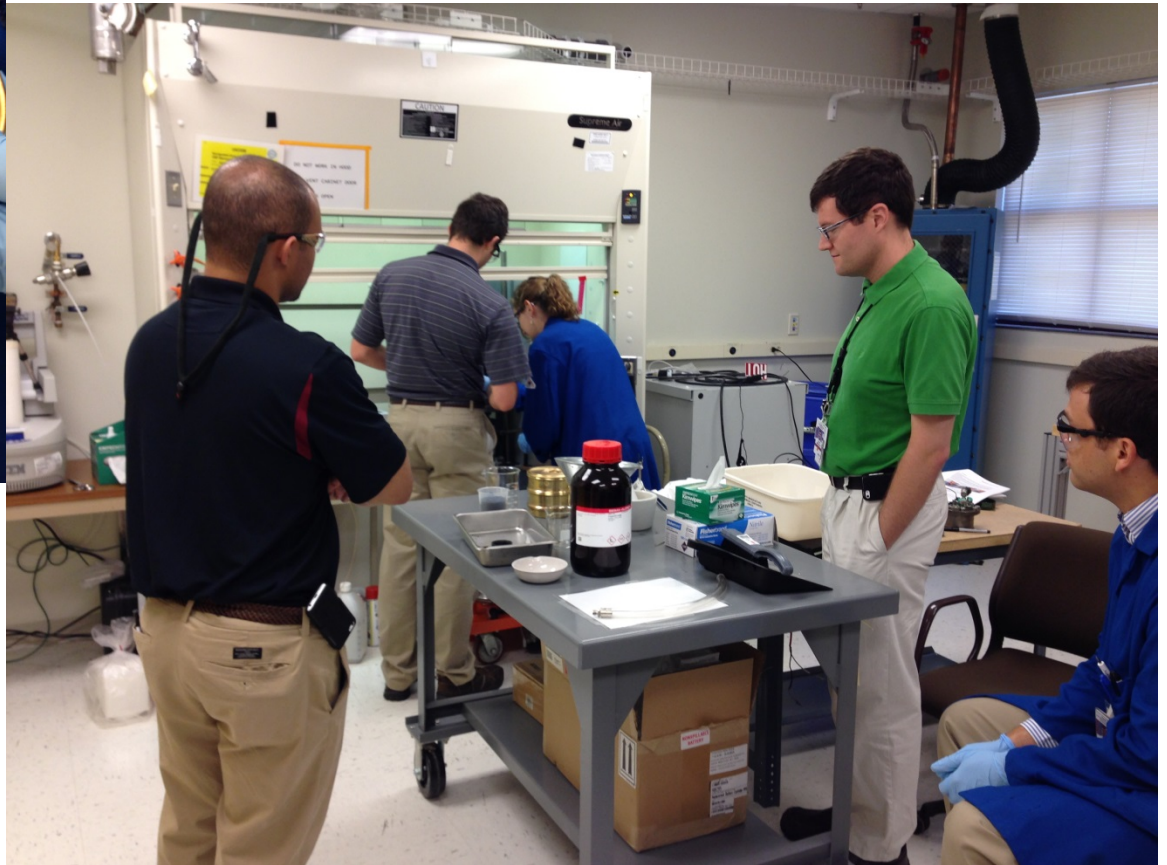


Fig. 2.—Reduction of copper oxide by hydrogen. Curve I, reduction at 150°. Curve II, reduction at 200°.

Robert N. Pgage, H. S. Taylor, "Reduction of Cupric Oxide by Hydrogen", J. Am. Chem. Soc., 1921, 43 (10), pp 2179

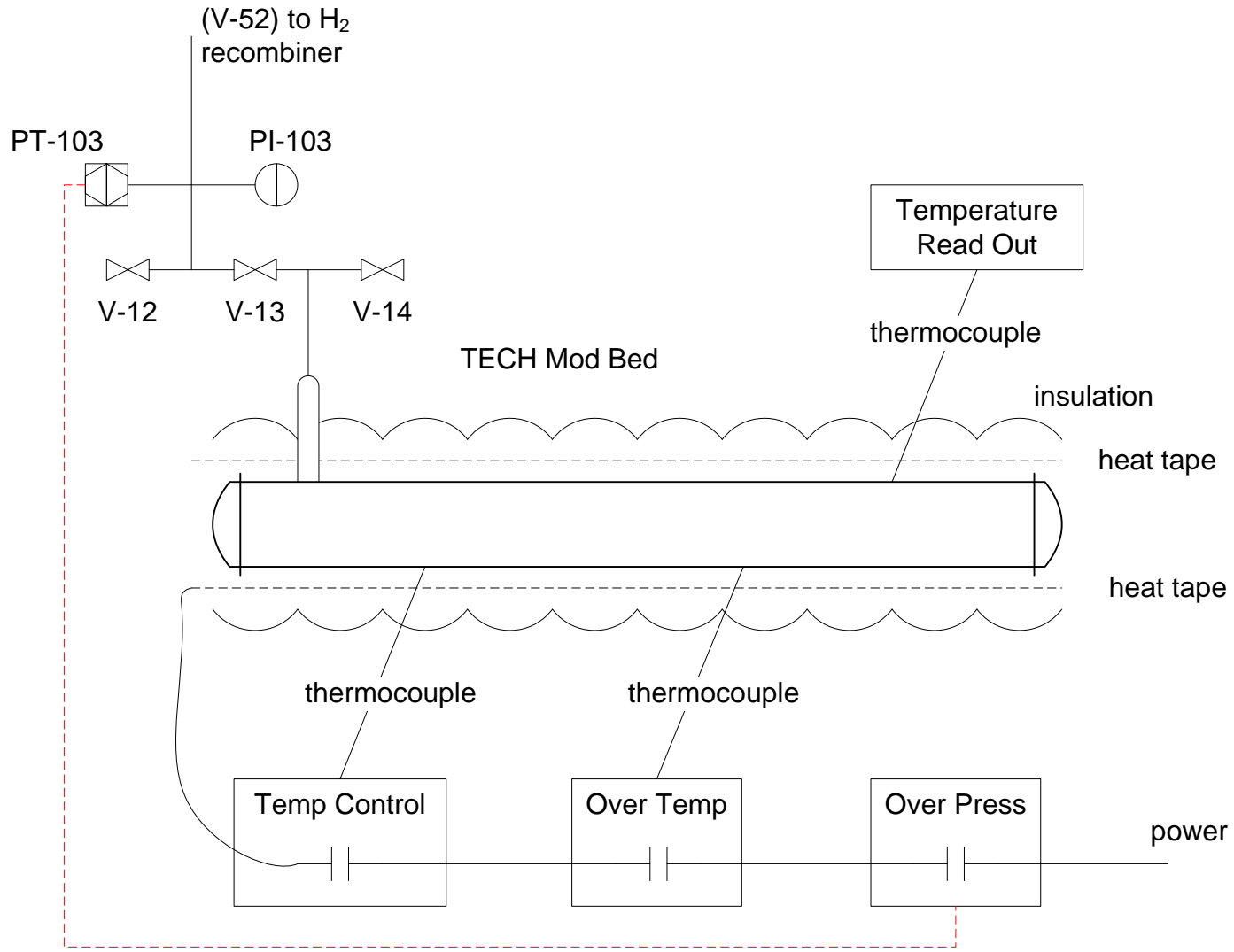


# Preparation of Copper Oxide

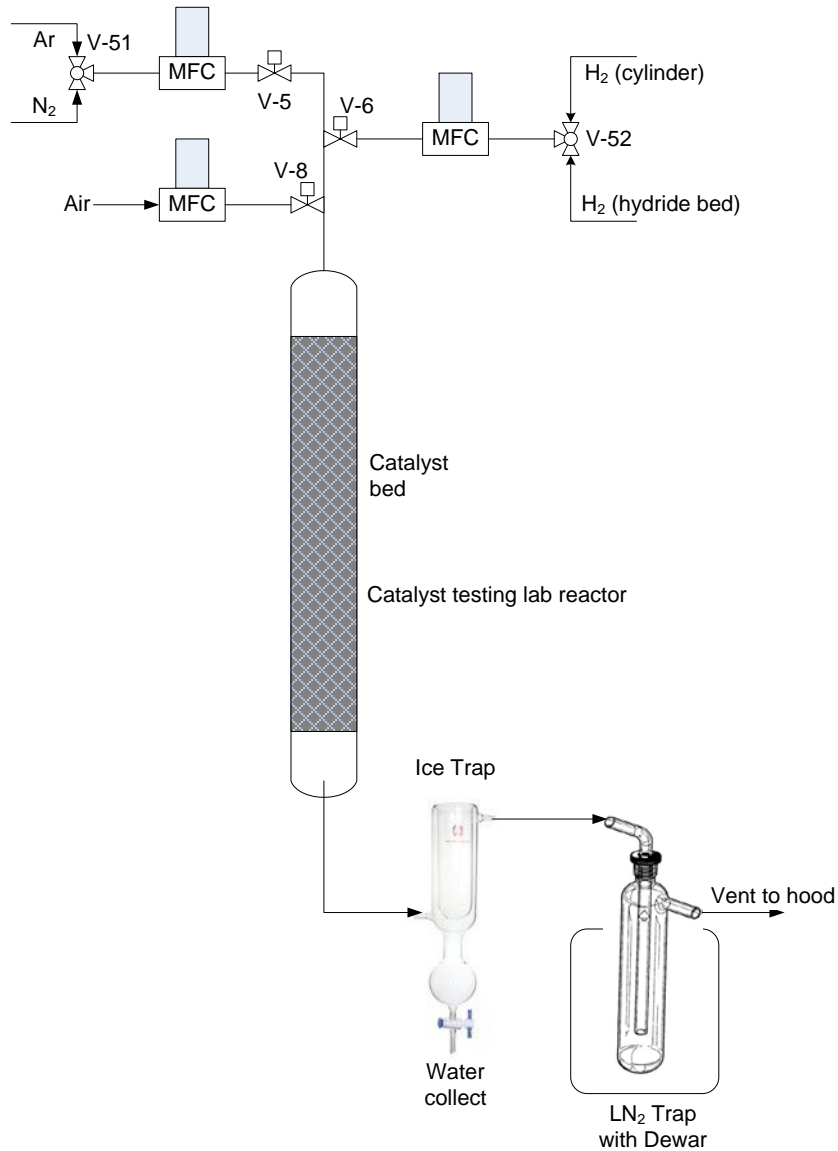




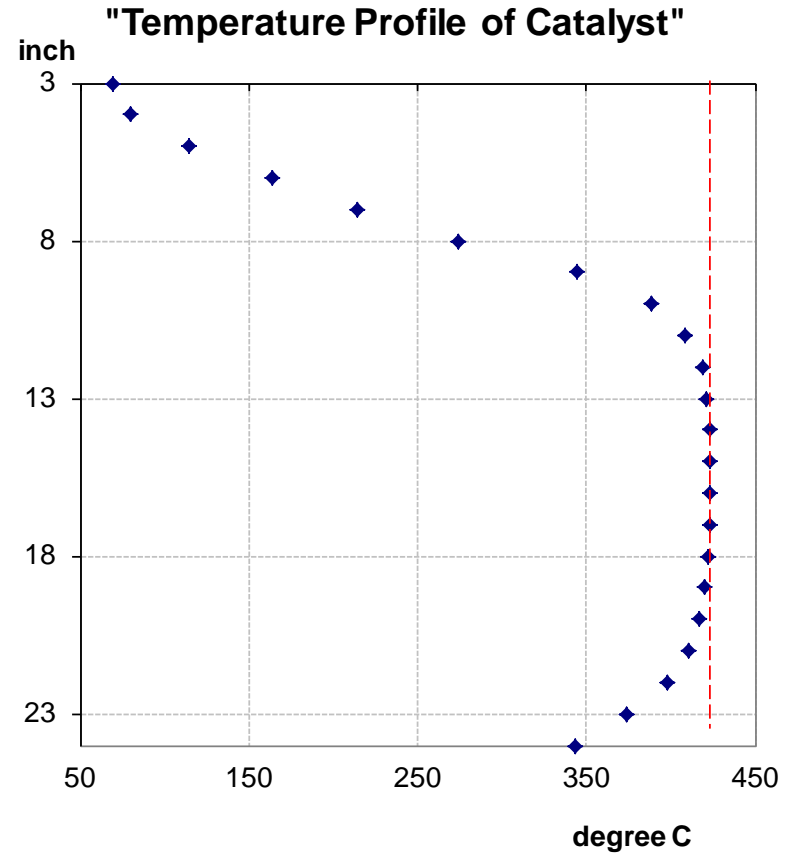
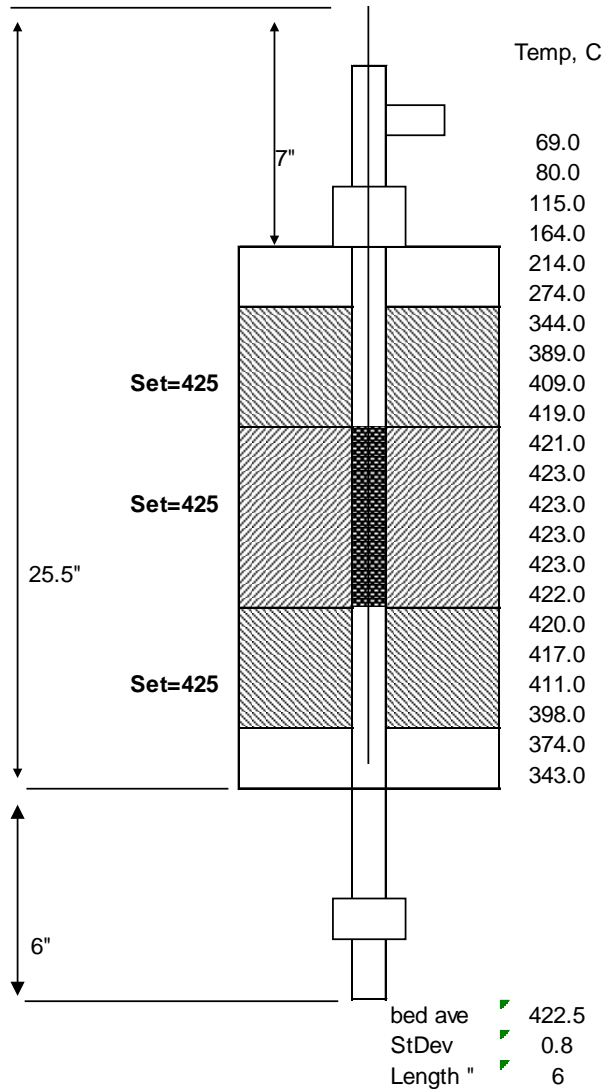
# Hydride Bed Setup



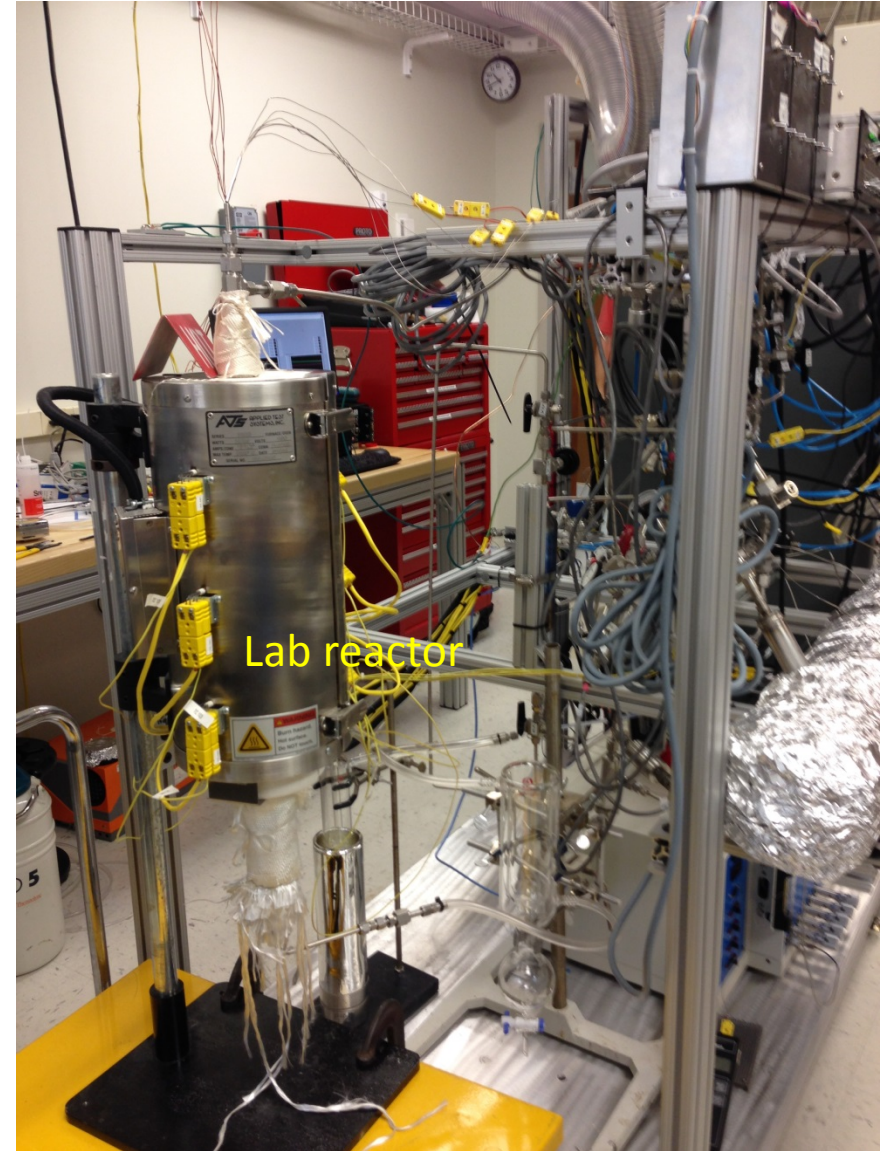
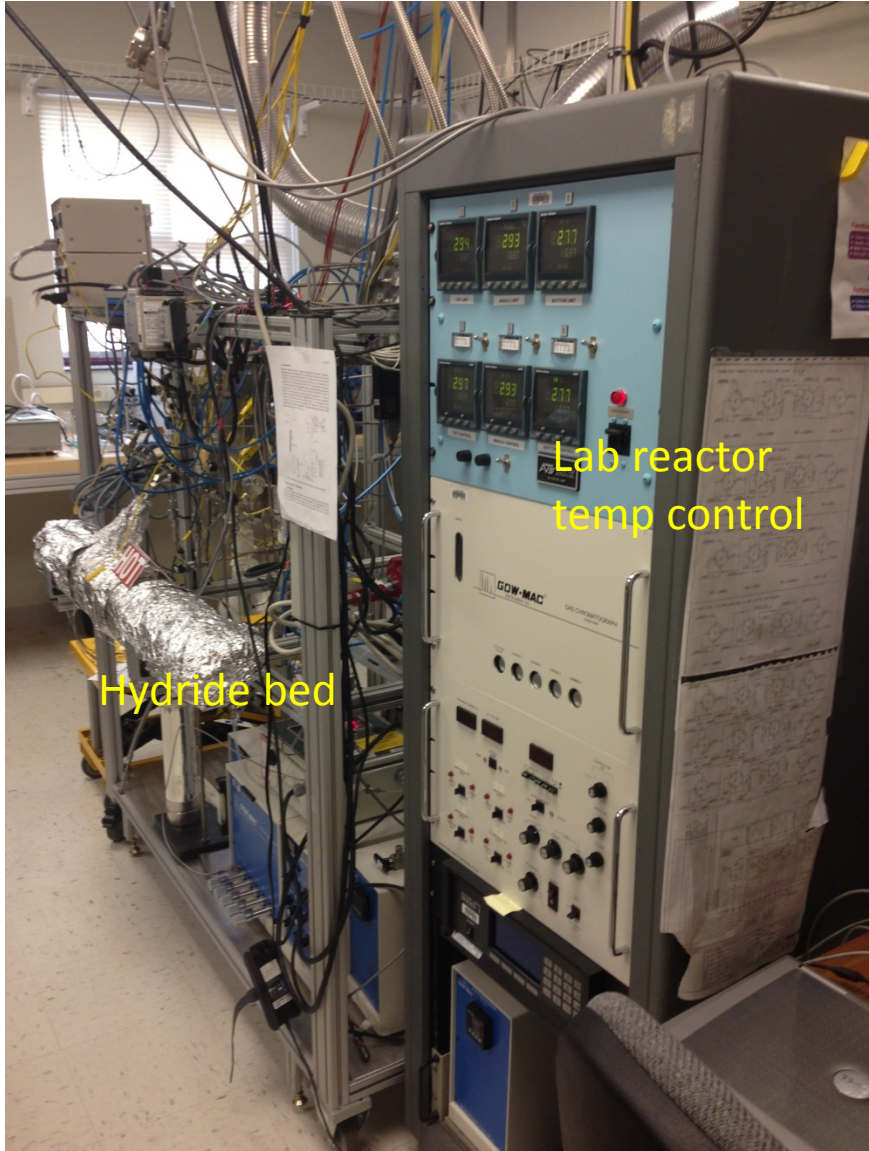
# Catalyst Bed Setup



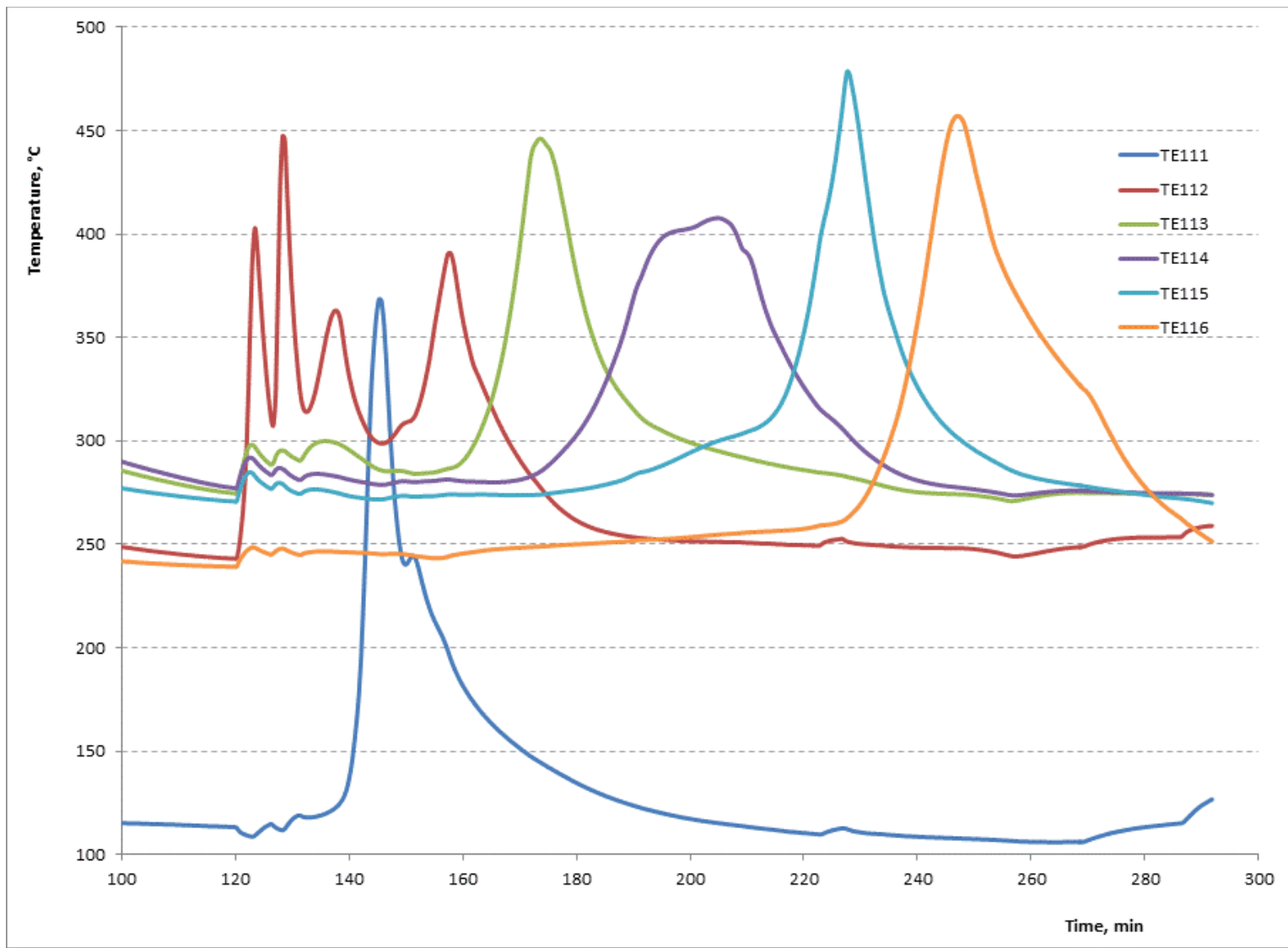
# Typical Tube Reactor Temperature Profile



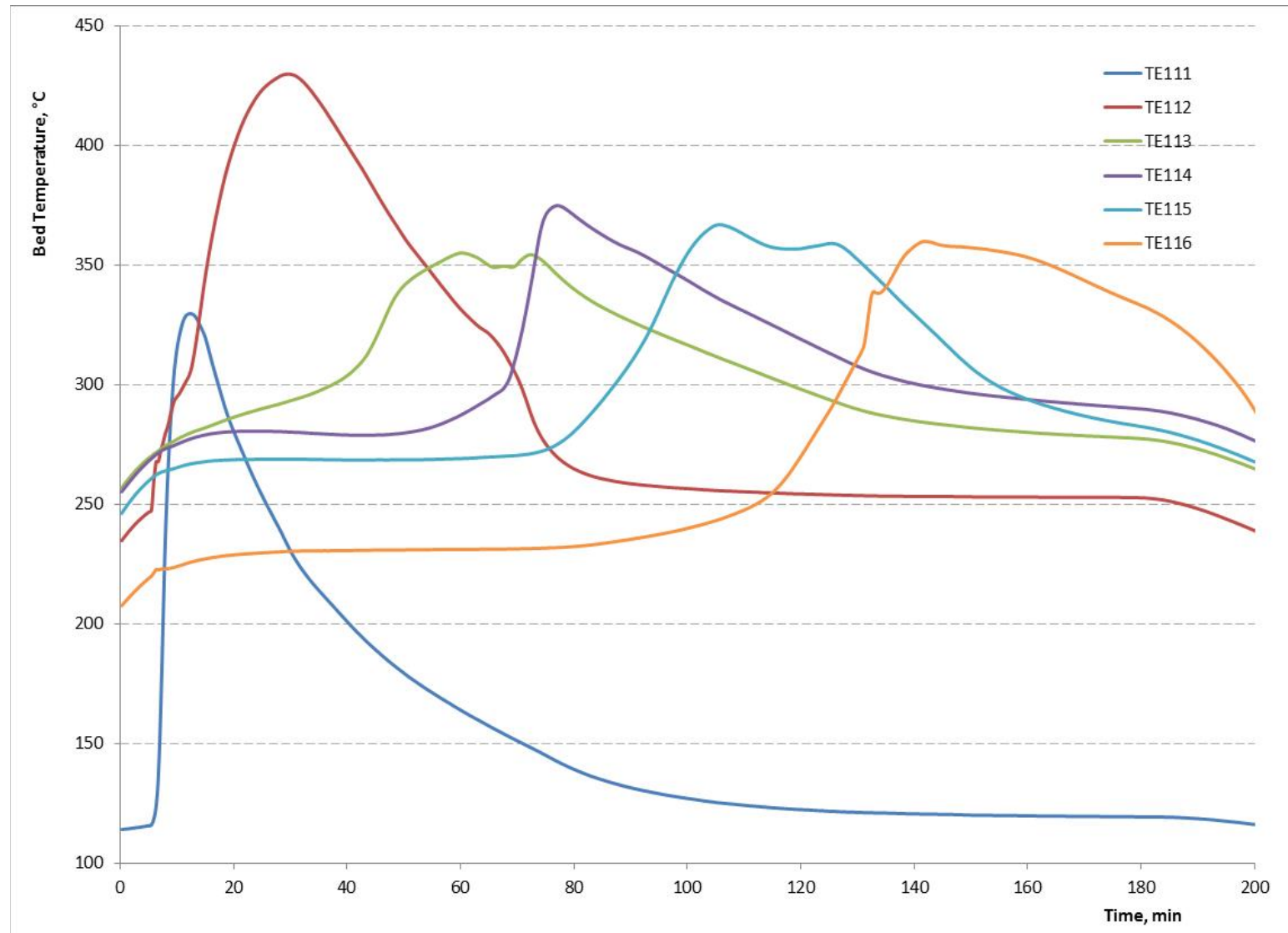
# Hydrogen Oxidation in Progress



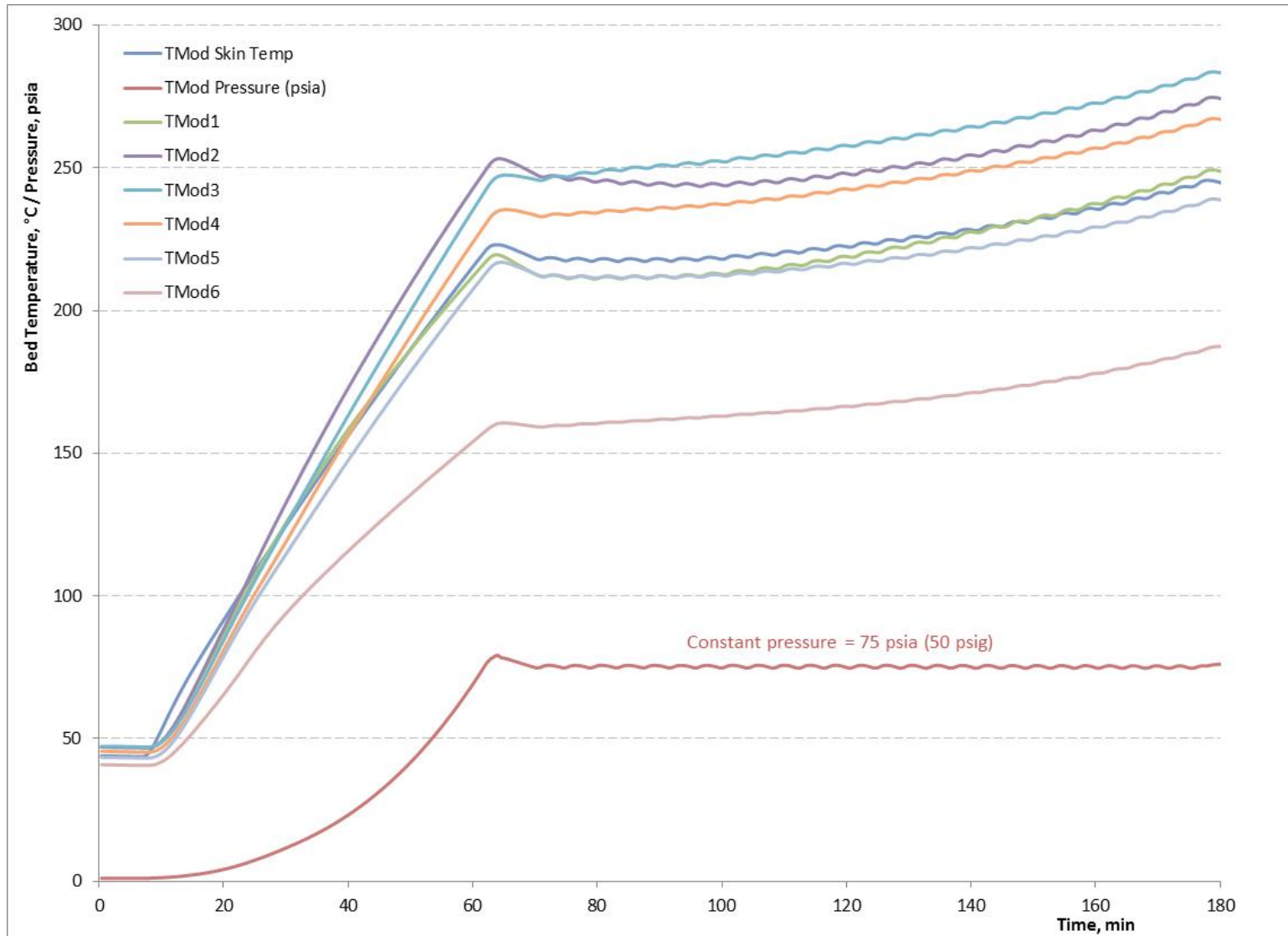
# Exothermic Fronts during Hydrogen Oxidation



# Exothermic Fronts during Copper Bed Regeneration



# Hydride Bed Temperature and Constant Pressure Control



# TCAP Raffinate Purity: 99.998% H<sub>2</sub>

H <sub>2</sub> + CuO → H <sub>2</sub> O + Cu	Water D/H Ratio, ppm
Test 1	23.2
Test 2	20.3
Test 3	20.5
Test 4	20.1

Single hydride bed supplied gas for 4 tests:

- D<sub>2</sub> reduction ratio: 25000 (50% in feed to 20 ppm)
- Results could be better if residue heel in hydride were purged clean
- Oxidation and bed regeneration can combine with co-feed H<sub>2</sub> and air

*Acknowledgements: Jim Klein for suggesting copper oxide for hydrogen oxidation; Katie Heroux for providing TechMod hydride bed; Greg Staack, Ray Battles, Kurt Breiting, Sherolyn Bishop, Robert Plonski and Joe Wheeler for reviewing safety documentations; Sharon Redd and Donna Allison for lab supports.*